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## Mole Conversions Name Chem Worksheet 11 3

**mole conversions name chem worksheet 11-3** - the mole is a unit for count, as is the dozen. a dozen is 12 items, but a mole is 602 000 000 000 000 000 000 000 , or  $6.02 \times 10^{23}$  particles. using the periodic table we can find the molar mass, or the mass of a mole of a substance. the molar volume is the volume of a mole of substance. all gases have the same molar volume when measured at **name mole conversions - doccasagrande** - wks name mole conversions period part i: find the molar masses (mm) for the following substances.(look up the mass of each element on the periodic table and add them all up.) write all molar masses with at least 4 sig figs. **mole conversions worksheet - gorechem.weebly** - mole-particle conversions (use avogadro's number for your conversions) 1. how many moles of magnesium are in  $223.01 \times 10$  atoms of magnesium? 2. how many molecules are there in 4.00 moles of glucose,  $C_6H_{12}O_6$ ? 3. how many moles are  $1.20 \times 10^{25}$  formula units of calcium iodide? 4. how many formula units are in 12.5 moles of calcium phosphate? **mole calculations multiple choice review psi chemistry name** - mole calculations multiple choice review psi chemistry name \_\_\_\_ the mole and avogadro's number 1) what is the si unit for measurement of number of particles in a ... for which of the following conversions does the value of the conversion factor depend **chemistry worksheet name: mole conversions and percent ...** - chemistry worksheet name: \_\_\_\_ mole conversions and percent composition block: \_\_\_\_ 5. answer the following questions about carbon dioxide,  $CO_2$  2. show your work, and round answers to the ones place. a. determine the molar mass of carbon dioxide. b. determine the % by mass of carbon in carbon dioxide. **mole to grams, grams to moles conversions worksheet** - mole to grams, grams to moles conversions worksheet what are the molecular weights of the following compounds? 1)  $NaOH$  2)  $H_3PO_4$  3)  $H_2O$  4)  $Mn_2Se_7$  5)  $MgCl_2$  6)  $NH_4_2SO_4$  there are three definitions (equalities) of mole. they are: 1 mole =  $6.02 \times 10^{23}$  particles 1 mole = molar mass (could be atomic mass from periodic table or molecular mass) **mole worksheet 2 - wappingers central school district** - name \_\_\_\_ period \_\_\_\_ mole worksheet #2 make the following conversions using unit analysis. use a separate piece of paper, show all work, and circle your final answer. (attach this sheet to your work). set a: one step problems: convert to moles: 1.  $12.04 \times 10^{23}$  atoms  $He$  2.  $3.01 \times 10^{23}$  atoms  $Cu$  3. **1 mole =  $6.02 \times 10^{23}$  atoms 1 mole = atomic mass (g)** - 1 mole =  $6.02 \times 10^{23}$  atoms 1 mole = atomic mass (g) 5. convert 9500g of iron to number of atoms in the sample. 9500 g  $Fe$  1 mole  $6.02 \times 10^{23}$  atoms =  $1.0 \times 10^{26}$  atoms 55.8g 1 mole 6. what is the mass of 0.250 moles of aluminum? **mole calculation worksheet - sheffieldschools** - 1 mole = molar mass (could be atomic mass from periodic table or molecular mass) 1 mole = 22.4 l of a gas at stp (you do not need to worry about this yet) each definition can be written as a set of two conversion factors. they are: 1 mole = molar mass(g) can be written as  $\frac{1 \text{ mole}}{\text{molar mass (g)}}$  **worksheet: mole problems name** - worksheet: mole problems name \_\_\_\_ chemistry: a study of matter © 2004, gpb 7.5 part 1: molar mass use the periodic table to find the molar masses of the following. **mass moles particles conversion - academic computer center** - 1 mole  $CaCO_3$  x 2 mole  $HCl$  / 1 mole  $CaCO_3$  = 2 mole  $HCl$  moles of substance a moles of substance b mole ratio this comes from the balanced chemical reaction! ...and it relates substances in the reaction! mole-to-mole conversions  $2HCl + CaCO_3 \rightarrow CaCl_2 + H_2O$  how many moles of  $HCl$  react with a mole of  $CaCO_3$ ? **name: mrs. piersa gfm, percent composition and mole ...** - name: mrs. piersa gfm, percent composition and mole conversions practice a) 112 g/mol b) 121 g/mol c) 149 g/mol d) 242 g/mol 1. what is the gram-formula mass of  $(NH_4)_3PO_4$ ? a) 1.0 g b) 4.8 g c) 48 g d) 480 g 2e gram-formula mass of a compound is 48 grams. **workshop 5: name mole conversions section** - when you are given a problem that involves conversions among moles, grams, number of particles, it is helpful to analyze the problem, write a road-map, consider what factors are available, set up the problem(s) using dimensions, put in the numbers, and then do the arithmetic. recall: 1 mole =  $6.022 \times 10^{23}$  particles (i.e. atoms or molecules or ...

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